

Controlling Magnetic Anisotropy in a Zero-Dimensional $S = 1$ Magnet Using Isotropic Cation Substitution

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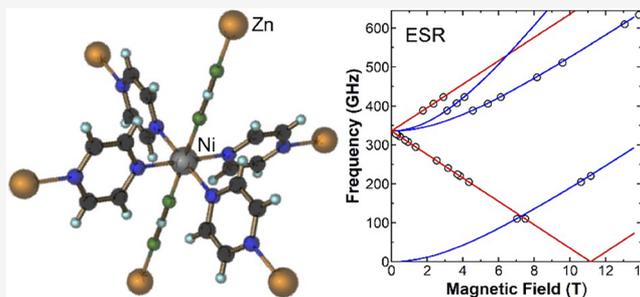
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ABSTRACT: The $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ($x = 0.2$; pyz = pyrazine) solid solution exhibits a zero-field splitting (D) that is 22% larger [$D = 16.2(2)$ K ($11.3(2)$ cm^{-1})] than that observed in the $x = 1$ material [$D = 13.3(1)$ K ($9.2(1)$ cm^{-1})]. The substantial change in D is accomplished by an anisotropic lattice expansion in the MN_4 ($M = \text{Zn}$ or Ni) plane, wherein the increased concentration of isotropic $\text{Zn}(\text{II})$ ions induces a nonlinear variation in M -F and M -N bond lengths. In this, we exploit the relative donor atom hardness, where M -F and M -N form strong ionic and weak coordinate covalent bonds, respectively, the latter being more sensitive to substitution of Ni by the slightly larger $\text{Zn}(\text{II})$ ion. In this way, we are able to tune the single-ion anisotropy of a magnetic lattice site by Zn -substitution on nearby sites. This effect has possible applications in the field of single-ion magnets and the design of other molecule-based magnetic systems.



INTRODUCTION

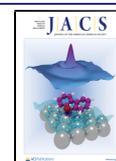
The idea of symmetry breaking, in which a system crossing a critical point gravitates to a lower symmetry state, was first noted in studies of rotating incompressible fluids under hydrostatic and gravitational equilibrium by Liouville¹ and Jacobi² in 1834. Since then, the concept has remained at the forefront of many areas of fundamental science, for instance, in theories of the early universe³ and the role of the Higgs boson in particle physics,⁴ catastrophe theory,⁵ structural phase transitions in liquids and solids,⁶ and magnetism.⁷ Recently, in various quantum materials, the observation of nematic electronic phases, in which the electrical resistivity of a previously highly symmetric material suddenly becomes anisotropic, and *potential control* of this effect, has excited a great deal of interest.⁸ While the above are examples of spontaneous symmetry breaking, here, we highlight a related effect driven by *chemical intervention* in which an anisotropic change in magnetic properties can be induced (and potentially controlled) by ionic substitution of an isotropic species.

We discuss the metal–organic framework (MOF) $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$, where pyz is pyrazine, and show that the replacement of $\text{Ni}(\text{II})$ by isotropic $\text{Zn}(\text{II})$ ions in the solid solution results in an anisotropic expansion of the square planes hosting the metal ions, leading to a modification of the magnetic anisotropy of the $\text{Ni}(\text{II})$ spins. Our study of the solid solution is motivated by measurements^{9–11} on the parent MOF $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ and the fact that a variety of desirable property enhancements in MOFs have been made

by introducing defects or disorder.^{12–14} Despite its simple structural framework, $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ exhibits complex magnetic behavior¹⁰ due to the interplay between the antiferromagnetic exchange interactions (J – mediated by HF_2^- , and J' – mediated by pyz ligands) and zero-field splitting (D). To help untangle the roles of D , J , and J' , $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ is synthesized and crystallized¹⁵ for $x = 0$ (diamagnetic host lattice) and $x = 0.2$ (zero-dimensional $S = 1$ magnet); each is found to be isomorphous to the $x = 1$ three-dimensional $S = 1$ magnet. Notably, electron-spin resonance (ESR) reveals a 22% increase in D from 13.3(1) K [$9.2(1)$ cm^{-1}] to 16.2(2) K [$11.3(2)$ cm^{-1}] as x decreases from 1.0 to 0.2. This is caused by the above-mentioned nonlinear variation in M -F and M -N bond lengths (where $M = \text{Zn}$ or Ni), resulting in anisotropic expansion of the MN_4 square planes. We attribute this behavior to softer pyz-N coordination that makes the M -N bonds more sensitive to Zn -substitution in stark contrast to the stronger ionic M -F bonds that change very little. This mechanism for locally tuning the ligand-field of NiN_4F_2 , via Zn -substitution of Ni on other lattice sites, represents a new resource in the “tool kit”¹¹

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employed to design and construct bespoke molecular magnets with desirable combinations of J , J' , and D . The effect could have particular significance in the field of single-ion magnets. The focus of research in this area has shifted in recent years to encompass molecule-based magnets constructed using d -block elements.^{16–19} However, we note that the tuning mechanism we report has wide applicability and could be employed in lanthanide complexes. In addition, the changes we observed can have ramifications in a number of different fields in which dilution is used to suppress interactions, identify functionally active sites, and enhance the effectiveness of single-ion probes such as ESR, and where the additional impacts of the dilution are not always considered.^{20,21}

EXPERIMENTAL SECTION

Synthesis and crystallization details for $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ($x = 0, 0.2, 1.0$) are described in the Supporting Information.¹⁵ The crystal structures of $[\text{M}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ($M = \text{Ni}, \text{Zn}$) are determined by single-crystal X-ray diffraction, while that of $M = \text{Zn}_{0.8}\text{Ni}_{0.2}$ is determined by Rietveld refinement of synchrotron powder-diffraction data. X-ray fluorescence spectroscopy is used to confirm the Zn/Ni ratio in the mixed-metal compound. The low-field magnetization M and susceptibility χ of the samples are measured using a commercial SQUID magnetometer, while high-field measurements are performed at the National High Magnetic Field Laboratory Pulsed-Field Facility using a 65 T short-pulse magnet.²² For the latter experiments, samples are tightly packed powders with masses of 1–2 mg. Data are taken in pulsed fields of up to 30 T and down to a base temperature of 0.5 K. High-field, high-frequency ESR spectra of powdered samples are recorded on a spectrometer at the EMR facility of the National High Magnetic Field Laboratory. Full details of all experiments are found in the Supporting Information.¹⁵

RESULTS AND DISCUSSION

Site Disorder in $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$. In our previous study,¹⁰ the complex magnetic behavior of $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ and a lack of suitably sized single crystals necessitated a multitechnique approach in which we found that the values of the Hamiltonian parameters D , J , and J' could only be determined precisely using inelastic neutron scattering. A primary aim of the current study is therefore to focus on D , removing complications due to J and J' by randomly diluting Ni(II) ions on an equivalent tetragonal lattice, thereby establishing a zero-dimensional $S = 1$ magnetic framework in which D can be measured unambiguously using electron spin resonance (ESR).

Hence, percolation calculations are performed to identify a Zn-substitution level sufficient to diminish the effect of exchange coupling between Ni(II) ions. This also allows the deduction of the minimum number of sites that must be occupied by Ni(II) ions on the Ni/Zn sublattice for a cooperative magnetic phase to emerge. A simple cubic lattice, spanning a hundred unit cells in each of the three Cartesian directions, is used as a model.

While a simple cubic lattice does not represent the true symmetry of the solid solution, it is functionally equivalent to the tetragonal Ni/Zn sublattice (see following section) for site-filling purposes. The lattice sites are randomly occupied, representing Ni on the site, or unoccupied, representing Zn on the lattice site, such that x (the concentration of Ni(II) ions) is the probability that a given site is occupied. This random distribution of Ni and Zn ions does not favor any specific bonding, nor the formation of specific clusters, and as such any Ni or Zn clustering is purely statistical. The percolation

threshold, the value of x for which there is a single cluster that spans from one face of the cubic lattice to an opposite face, is calculated to be $x = 0.311$, in perfect agreement with ref 23. (See Figure 1.) This demonstrates that $x = 0.2$, on which our

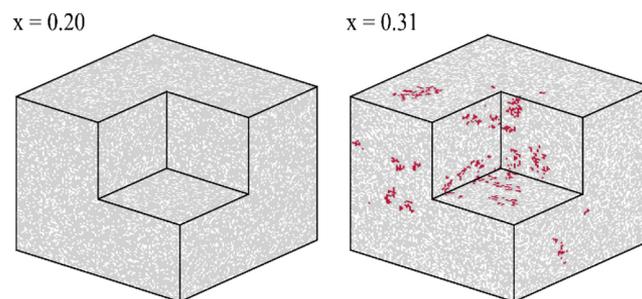


Figure 1. 3D illustrations of simple cubic lattices of face-sharing voxels for the Ni(II) concentrations $x = 0.2$ and 0.31 . Only voxels that intersect the surface are shown, and an eighth of the cubic lattice has been removed to reveal the interior. The largest ordered Ni cluster is shown in red, all other sites occupied by Ni are given in white and sites occupied by Zn are shown in gray.

experiments were performed, represents a safe margin below the percolation threshold; only small clusters of exchange-correlated Ni(II) ions will exist and hence no long-range magnetic order is expected. As shown below, the presence of small correlated clusters in the material does not adversely affect the determination of the zero-field splitting of the Ni(II) ion.

Crystal Structures. The three compounds ($[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$, $x = 0, 0.2, 1.0$) are isomorphous and crystallize in the $P4/nmm$ space group (Table 1). Metal cations occupy $42m$ sites and lie at the center of slightly compressed D_{4h} symmetric MN_4F_2 pseudo-octahedra. The coordinated pyz ligands adopt a 4 symmetry about M, each making an alternating dihedral angle $\approx \pm 73^\circ$. As shown in Figure 2, the structures consist of a cationic MOF-like $[\text{M}(\text{HF}_2)(\text{pyz})_2]^+$ framework, self-assembled from one-dimensional M-FHF-M linear chains that lie along the c -axis. Within the ab -plane, pyz ligands fuse the chains together to give two-dimensional square lattices. Charge neutrality is ensured by SbF_6^- ions that occupy pseudo-body-centered sites. These moieties are fixed in place by weak pyz C–H \cdots F interactions where H \cdots F ranges between 2.523 Å ($M = \text{Ni}$) to 2.576 Å ($M = \text{Zn}$) with the solid solution being intermediate at 2.559 Å.

A comparison of axial M-F and equatorial M-N bond lengths reveals key differences (Table 2 and Figure 3). First, the M-F bonds are slightly shorter than the M-N bonds. Second, the M-F bond length changes by 1.5% across the series, indicating the strong ionicity of the interaction. Third, we observe an overall greater variation of 2.9% among M-N coordinate bonds for different M, which steadily increase in length from 2.113(3) Å ($M = \text{Ni}$) to 2.143(3) Å ($M = \text{Zn}_{0.8}\text{Ni}_{0.2}$) to 2.174(3) Å ($M = \text{Zn}$).

The growth in the anisotropy ratio $[d(\text{M-N})/d(\text{M-F})]$ with increasing Zn content accelerates markedly after 80% Zn-doping (Figure 3), despite most of the lattice sites already being occupied by isotropic Zn(II) ions; the latter have a 7.2% larger ionic radii than Ni(II).²⁴ To the best of our knowledge, such behavior and its impact on magnetic properties (see below) have not been previously reported.

Table 1. X-ray Structural and Refinement Parameters for $[M(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ($M = \text{Ni}, \text{Zn}$ and $\text{Zn}_{0.8}\text{Ni}_{0.2}$) Determined at 298 K

Parameter	M = Ni	M = Zn	M = $\text{Zn}_{0.8}\text{Ni}_{0.2}$
Method	Single xtl	Single xtl	Powder
Instrument	Bruker d8	Bruker d8	APS 11-bm
λ (Å)	1.54178	1.54178	0.41275
Emp. formula	$\text{C}_8\text{H}_9\text{F}_8\text{N}_4\text{NiSb}$	$\text{C}_8\text{H}_9\text{F}_8\text{N}_4\text{ZnSb}$	$\text{C}_8\text{H}_9\text{F}_8\text{N}_4\text{Zn}_{0.8}\text{Ni}_{0.2}\text{Sb}$
Mol Wt (g/mol)	493.65	500.31	498.97
Crystal class	tetragonal	tetragonal	tetragonal
Space group	$P4/nmm$	$P4/nmm$	$P4/nmm$
$a = b$ (Å)	9.9430(5)	10.0871(2)	10.05603(1)
c (Å)	6.4521(5)	6.4847(3)	6.477771(1)
V (Å ³)	637.88(8)	659.82(4)	655.056(2)
Z	2	2	2
ρ_{exp} (g/cm ³)	2.570	2.518	2.533
GOF or χ^2	1.211	1.201	10.29
R_1^a or R_{exp}^a	0.0402	0.0293	0.0591
wR_2^a or R_{wp}^a	0.0926	0.0686	0.1345

^aEquations defining these parameters can be found in the deposited CIFs.

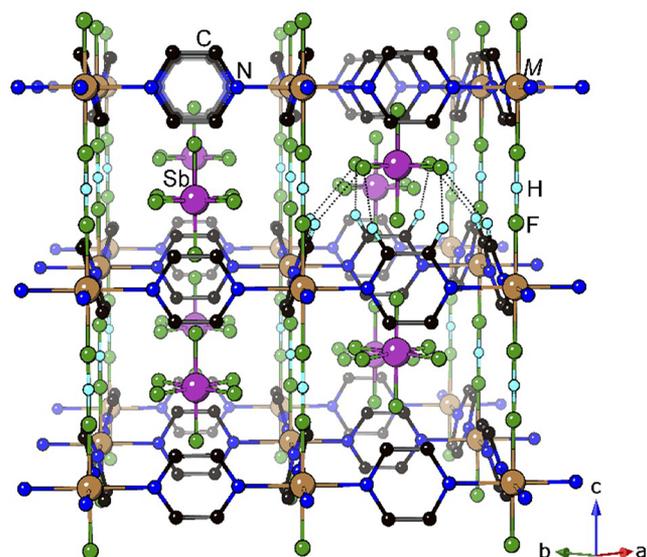


Figure 2. Representative 298 K crystal structure of $[M(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ along with the atom labeling scheme. For the $M = \text{Zn}_{0.8}\text{Ni}_{0.2}$ solid solution, the cations are randomly distributed over the M sites. Dotted lines represent weak C–H...F interactions between pyz and SbF_6 ; most pyz H atoms are omitted for clarity.

Table 2. Selected Bond Lengths (Å) and Bond Angles (deg) for $[M(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ($M = \text{Ni}, \text{Zn}$ and $\text{Zn}_{0.8}\text{Ni}_{0.2}$) Obtained at 298 K

Bond length or angle	M = Ni	M = Zn	M = $\text{Zn}_{0.8}\text{Ni}_{0.2}$
M-F (Å)	2.078(4)	2.109(3)	2.104(3)
M-N (Å)	2.113(3)	2.174(3)	2.143(3)
H-F (Å)	1.148(4)	1.133(3)	1.131(3)
F...F (Å)	2.296(4)	2.266(3)	2.260(3)
M-FHF-M (Å)	6.4521(5)	6.4847(3)	6.477771(1)
M-pyz-M (Å)	7.0308(5)	7.1327(2)	7.11069(1)
M-pyz dihedral (deg)	72.9	73.3	73.4
M-FHF-M (deg)	180	180	180
M-pyz-M (deg)	180	180	180

Bulk Magnetometry. SQUID magnetometer data¹⁵ on powder samples of $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ are shown as

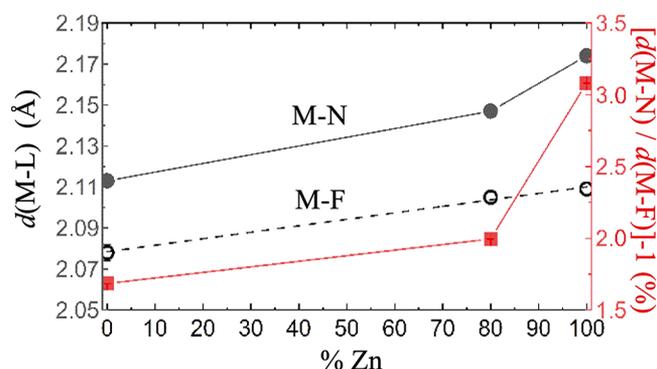


Figure 3. Plots of M-N and M-F bond lengths (Å) and MN_4F_2 anisotropy ratio vs %Zn composition. Lines are guides to the eye only. The closed/open circles, along with solid and dashed black lines, correspond to the left axis whereas the red squares and red line correspond to the right axis.

$\chi(T)$ ($= M/H, \mu_0 H = 0.1$ T) in the main plot of Figure 4 along with the previously published data for the pure Ni(II)

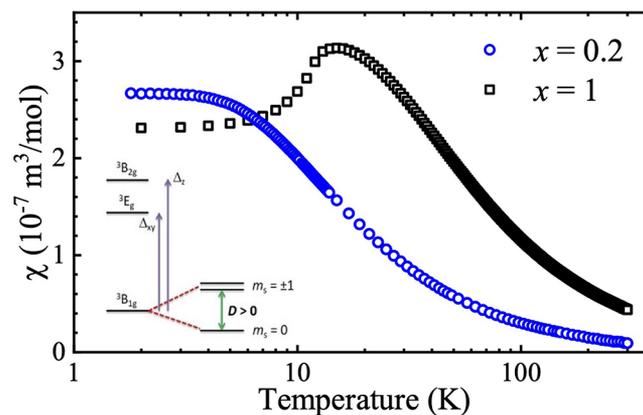


Figure 4. Susceptibility data on a log- T scale for powdered samples of $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$, $x = 0.2$ (blue circles) and $x = 1$ (black squares) taken in applied fields of 0.1 T. (Inset) Illustration showing the zero-field energy-level diagram and the splitting of the Ni(II) ground-state.

compound.¹⁰ The data for the parent material show the characteristic form of a quasi-one-dimensional $S = 1$ antiferromagnet. The rapid decrease in the data on being cooled below about 12 K coincides with the three-dimensional long-range ordering temperature, as confirmed by heat-capacity and neutron-diffraction measurements.¹⁰ The effect of Zn-substitution is evidenced by the data for the $x = 0.2$ material, which are reminiscent of expectations for a polycrystalline $S = 1$ material with small or negligible exchange interactions (see, e.g., ref 29). No evidence for a transition to long-range magnetic order is apparent, as predicted by the percolation calculations.

Below about 4 K, the $\chi(T)$ data for $x = 0.2$ reach a plateau because of the zero-field splitting of the $^3B_{1g}$ Ni(II) ground-state in a tetragonal ligand field.^{25,26} While the orbital-angular momentum is largely quenched, the spin-orbit coupling perturbs the magnetic m_s states as illustrated in the inset of Figure 4.²⁷ On the basis of spectrochemical predictions, the $m_s = 0$ spin singlet is lower in energy owing to a weaker axial ligand-field. The Δ_{xy} and Δ_z excitations arise from the respective difference in the equatorial Ni–N and axial Ni–F ligand-fields, which lead to the observed magnetic anisotropy.²⁸

Pulsed-field magnetometry¹⁵ data ($T \approx 0.5$ K) for $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ and $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ are compared in Figure 5. The $M = \text{Ni}$ parent compound shows a

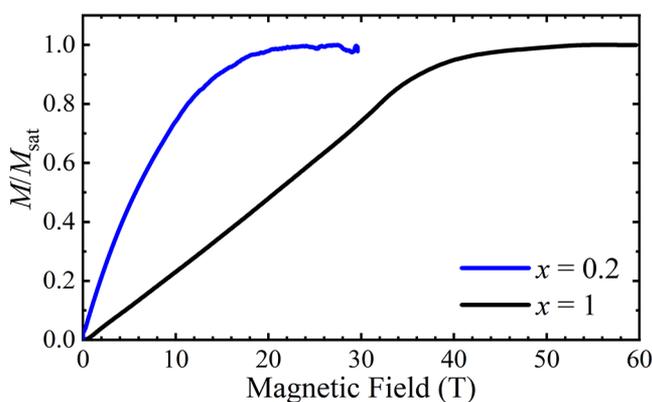


Figure 5. Plot of the magnetization normalized to the saturation value (M/M_{sat}) for powdered samples of $[\text{Zn}_{1-x}\text{Ni}_x(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$, $x = 0.2$ (blue line) and $x = 1$ (black line). The data for $x = 1$ are from ref 10. For the latter, the near-linear, slow rise in $M(H)$ prior to saturation is due to AFM correlations between Ni ions. For both, the observed rounded approach to M_{sat} is attributed to the presence of significant D values.

slow, near-linear rise in $M(H)$ prior to saturation that is the hallmark of AFM correlations between Ni ions,¹⁰ pointing to significant J values. By contrast, in the solid solution, this feature is absent, suggesting that the desired reduction of exchange interactions has been achieved. However, for both compounds, the rounded approach to M_{sat} is attributable to a large value of D .

The magnetization data for $x = 0.2$ are calibrated using low-field SQUID $M(H)$ results performed at a similar temperature, and the saturation magnetization at high fields was found to be $M_{\text{sat}} = x\tilde{g}S = 0.38(1)\mu_{\text{B}}/\text{formula unit}$ where \tilde{g} is the powder averaged g -factor. Using $\tilde{g} = 2.210(6)$ derived from ESR (see below), we found that $x = 0.17(1)$, in agreement with the value $[x = 0.20(5)]$ obtained from X-ray fluorescence.¹⁵

Even though the $x = 0.2$ material lies below the percolation threshold, it will still contain isolated regions of exchange-correlated Ni(II) ions, which will have an effect on the bulk magnetometry data and the reliability of any fits to these data using a simple exchange-free model. Instead, for an unambiguous measurement of the microscopic local environment around the Ni(II) ions, we turn to ESR.

Electron Spin Resonance. Efforts to interpret ESR data from powder samples of the pure $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ end-member failed owing to the energetic competition between D and J regardless of the temperature and field/frequency combination used in the experiment.¹⁰ However, good quality spectra were obtained in this work¹⁵ for $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$. Figure 6 shows 10 K ESR spectra measured and

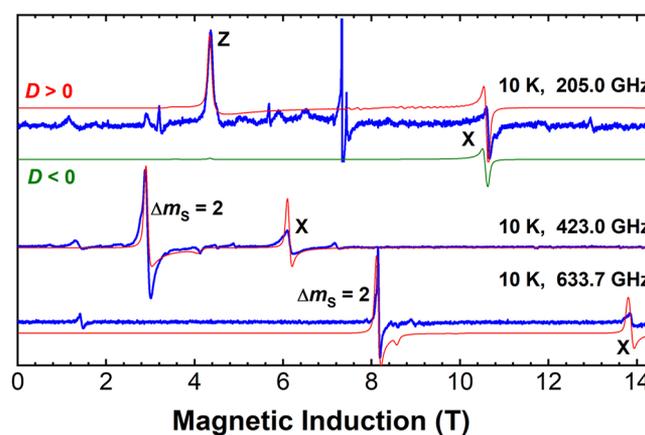


Figure 6. ESR spectra of $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ measured at 10 K for the indicated frequencies. Experimental data and simulations are shown as the blue and red lines, respectively. The green line corresponds to the same simulated parameters, but with a negative value of D [$D = -16.2(2)$ K].

simulated at 205, 423 and 633.7 GHz. As expected for a powdered $S = 1$ material, multiple resonances are observed between the split triplet energy levels. At a particular frequency and for the field applied along the x , y , or z -directions, two transitions that obey the ESR selection rule ($\Delta m_s = \pm 1$) are possible. A general good agreement is found between the significant features in the experimental data and simulations for the x - and z -orientations and the half-field $\Delta m_s = 2$ transition. Formally, the half-field transition is forbidden; however, the selection rule is relaxed when the Zeeman energy is comparable to the zero-field splitting. Additional small features in the spectra are not accounted for by the simple model and are likely attributable to the presence of small, exchange-correlated regions in the sample, as predicted by the percolation calculations. From the frequency dependences shown in Figure 7, the derived parameters were $g_x = g_y = 2.240(4)$, $g_z = 2.15(1)$, $D = 16.2(2)$ K [$11.3(2)$ cm^{-1}].

Anisotropic Lattice Strain. The structural results described above demonstrate that the materials synthesized in this study, $[\text{Zn}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ and the solid solution $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$, are isomorphous to $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ ^{9–11} as well as the Co- and Cu-analogs.³⁰ In contrast to the parent $[\text{Ni}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$ material, the solid solution shows no evidence of long-range magnetic correlations and instead resembles an array of zero-dimensional $S = 1$ anisotropic magnets. Moreover, we obtain a 22% increase of D

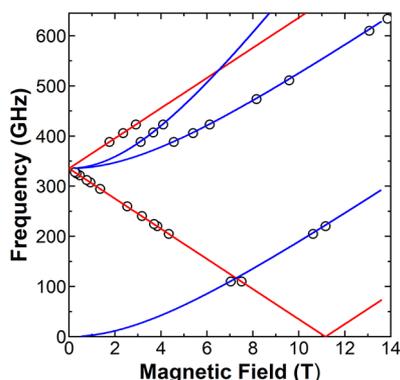


Figure 7. Frequency-dependencies of the ESR resonance fields observed in $[\text{Zn}_{0.8}\text{Ni}_{0.2}(\text{HF}_2)(\text{pyz})_2]\text{SbF}_6$. Red and blue traces show resonances calculated at the molecular orientations z and x , respectively. Circles are experimental data.

from 13.3(1) ($M = \text{Ni}$) to 16.2(2) K ($M = \text{Zn}_{0.8}\text{Ni}_{0.2}$) (i.e., from 9.2(1) cm^{-1} to 11.3(2) cm^{-1}).

The enhancement in the axial magnetic anisotropy is brought about via an anisotropic lattice strain caused by replacing 80% of Ni(II) ions with the 7.2% larger Zn(II). The key to this effect is that although the substituted Zn(II) ion is isotropic, the organic framework itself supplies the necessary anisotropy. Taking into account the difference in donor-atom hardness, with F being harder than N, the effect exploits the dissimilar ionic M-F and coordinate covalent M-N bond strengths to achieve the enhancement in D ; weaker M-N bonds are more sensitive to substitution by the larger Zn(II) ion. Hence, the increased concentration of Zn causes a nonlinear change in M-N bond lengths compared to M-F and yields an anisotropic lattice expansion within the MN_4 plane. Remarkably, the comparatively small percentage change in bond lengths leads to the significant boost in D .

CONCLUSION

Our findings reveal a novel mechanism by which strain can be used to adjust the ligand-field of magnetic ions and hence their local magnetic anisotropy through the introduction of nonmagnetic ions on nearby lattice sites. We suggest that in this way, the combination of anisotropic organic frameworks and metal ions of different radii may be exploited as an approach to tune the single-ion properties of magnetically isolated spins. These results will be useful in the investigation and control of magnetization reversal and slow relaxation in self-assembled arrays of single-ion magnets^{16–19} composed of d - or f -elements for use in spin-based electronic devices.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/jacs.0c12516>.

Full experimental details, including syntheses, X-ray fluorescence, single crystal X-ray diffraction and data collection, synchrotron powder diffraction and Rietveld analysis, SQUID and pulsed-field magnetometry and electron-spin resonance (PDF)

Accession Codes

CCDC 2026750–2026752 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by

emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Notes

The authors declare no competing financial interest.

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